

# Bisphenol A Removal by Adsorption Using Waste Biomass: Isotherm and Kinetic Studies

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**Abstract:** BPA is categorized as a new emerging pollutant that mimics the structure and function of hormone estrogen for humans and animals. The removal of BPA had been concerned because human exposure toward BPA is widespread, and it is harmful to humans and the ecosystem. The aim of the study was to remove BPA by adsorption technique and understand the application of agricultural waste material as potential adsorbent. The increasing of the surface area of post-adsorption from 4.80 m<sup>2</sup>/g to 5.83 m<sup>2</sup>/g indicated that the sulphuric acid treatment responsible for the growth in the porosity of the banana bunches, increase the number of available binding sites and its ion-adsorption capacity. The highest removal of BPA by the banana bunch and coconut bunch was obtained in the following condition: the temperature at 25°, pH 3, dosage at 100 mg, and agitation at 150rpm. The adsorption process was well described with the Langmuir isotherm, while the best correlation with the kinetic study of BPA adsorption was the pseudo-second-order model. The negative values of  $\Delta H^\circ$  suggested that the adsorption of BPA onto banana bunch (-13.748 J mol<sup>-1</sup>) and coconut bunch (-5693.67 J mol<sup>-1</sup>) is exothermic in nature.

**Keywords:** bisphenol A; banana bunch; coconut bunch; isotherms; pseudo-second-order kinetics.

**Nomenclature:** C<sub>0</sub> - Initial BPA concentration in the liquid phase (mg/L); C<sub>e</sub> - Equilibrium of BPA concentration in the liquid phase (mg/L); C<sub>t</sub> - BPA concentration at time t in the liquid phase (mg/L); V - Solution volume (L); M - Mass of the dry adsorbent (g); q<sub>e</sub> - The amount of sorbed BPA at equilibrium (mg/g); q<sub>t</sub> - The amount of sorbed BPA at time t (mg/g); Q<sub>max</sub> - Maximum sorption capacity in Langmuir model (mg/g); K<sub>L</sub> - Langmuir constant (L/mg); K<sub>F</sub> - Freundlich constants associated with adsorption capacity; n - number of isotherm parameters; A - Temkin equilibrium binding constant (L/g); B - Temkin constant related to the heat of sorption (J mol<sup>-18</sup>); K<sub>1</sub> - Pseudo-first-order rate constant; K<sub>2</sub> - Pseudo-second order rate constant; K<sub>c</sub> - Concentration factor of a chemical equilibrium equation; R - Natural gas constant (m<sup>3</sup>·atm·K<sup>-1</sup>·mol<sup>-1</sup>).

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## 1. Introduction

BPA stands for bisphenol A that first discovered in the 1890s and industrialized to produce strong plastics and resins since the 1960s. Currently, BPA is categorized as a new emerging pollutant that mimics the structure and function of hormone estrogen for humans and

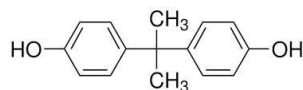
animals [1]. It is persistently being used in consumer goods and containers that store food and beverages, such as water bottles [2,3]. The removal of BPA had been concerned because human exposure toward BPA is widespread, and it is harmful to humans and the ecosystem. Exposure of BPA can cause a possible health effect on the brain and prostate of infants and children [1,4]. The abundance uses of BPA had led to the emission of toxicity of BPA into a water source and affect the aquatic life and contaminate the drinking water. BPA can enter the body through ingestion; especially, then, it will interfere with human hormonal function. The previous study revealed that BPA was detected at a dangerous level in the urine sample of children and adults [3]. For this study, the adsorption treatment had been proposed to eliminate the BPA in aqueous solution. This technique is commonly known as a cheap and efficient method, good removal ability, insensitivity, and simple design. For the removal of emerging pollutants by adsorption technique, commonly activated carbon (AC) was used as adsorbent [5-7]. However, currently, adsorbent from agricultural waste is most preferred and used widely compared to commercial activated carbon. Agricultural waste has been widely investigated as low-cost adsorbent since it is known to contain a variety of functional groups and basic components which have potential sorption capacities for several pollutants, other than environmentally friendly and reliable resources [8-11].

For this study, banana bunch and coconut bunch were used as adsorbent and treated with acid sulfuric for chemical modification. The agricultural waste came from different kinds of shape, color, and sources; thus, it needs pre-treatment before use as adsorbent. All colors, metals, and any soluble organic compounds of adsorbent ought to remove with different kinds of modifying agents, such as oxidizing agents, organic compounds, acid and base solutions, dyes, and mineral [12]. The objective of the study was to remove BPA by adsorption technique and understand the application of agricultural waste material as potential adsorbent. Various parameters condition on the adsorption and removal of BPA were also investigated.

## 2. Materials and Methods

### 2.1. Sorbate and chemicals.

BPA ( $C_{15}H_{16}O_2$ , molecular weight = 228.29,  $\lambda_{max}$  = 278 nm; purity 97% of Sigma-Aldrich Milwaukee, USA) was used as sorbates in this study. The stock solution of BPA was prepared by dissolving it (20 ppm) into 1 L distilled water. Figure 1 shows the molecular structure of BPA. Sulfuric acid (95-97%) was purchased from Merck (Germany). Ethanol (96%) and other chemicals were purchased from Qrec (New Zealand).



**Figure 1.** Molecular structure of BPA.

### 2.2. The sorbent.

Banana bunch (*Musa acuminata*) and coconut bunch (*Cocos nucifera*) were selected as the sorbent for the adsorption of BPA after selection. The sorbents were washed, cut (30 mesh sieve), oven-dried (105 °C), and pulverized before activation. The biomass was activated with  $H_2SO_4$  in ratio 1:1 (sorbent: acid) [13]. The dried adsorbent was washed with distilled water

until the pH becomes neutral. The treated sorbents and stored in an airtight container for subsequent use.

### 2.3. Batch experiment.

Sorption experiments were accomplished by agitating (150rpm) 100 mg of adsorbent with 20 ppm od BPA solution (pH 3) at room temperature. The blank experiment was performed by preparing a solution without adsorbent to ensure that no BPA was adsorbed onto the flask, which would interfere during spectrophotometer analysis. The various conditions were implemented to know the effect of some parameters in the sorption of BPA, as summarized in Table 1.

**Table 1.** Summarizes for all experiments of batch adsorption.

Parameters	Initial pH	Dosage (mg)	Temperature (°C)	Agitation (rpm)
Effect of Initial pH	3 - 9	100	25	150
Effect of Dosage	3	50 - 500	25	150
Effect of Temperature	3	100	25 - 60	150
Effect of Agitation	3	100	25	50 - 250

After a specific time interval, adsorbents were separated by filtration utilizing Whatman filter paper (125mm) No 1, and the initial and residual BPA concentration were determined using a (NANOCOLOR® VIS, Macherey-Nagel, Germany) at 278 nm. All experiments were performed in triplicate. The amount of BPA adsorbed onto banana bunch and coconut bunch at time  $t$ ,  $q_t$  (mg/g), and at equilibrium time  $q_e$  (mg/g) was calculated by equation (1), (2), and (3), respectively.

$$R (\%) = \frac{C_o - C_e}{C_o} \times 100 \tag{1}$$

$$q_t = \frac{(C_o - C_t)V}{M} \tag{2}$$

$$q_e = \frac{(C_o - C_e)V}{M} \tag{3}$$

### 2.4. Sorption isotherm model.

In this study, the Langmuir, Freundlich, and Temkin models were used to examine the sorption behavior of BPA onto the adsorbent, as calculated by equation (4), (5), (6), respectively.

$$q_e = \frac{Q_{\max} K_L C_e}{1 + K_L C_e} \tag{4}$$

$$q_e = K_F C_e^{1/n} \tag{5}$$

$$q_e = B \ln A + B \ln C_e \tag{6}$$

### 2.5. Sorption kinetic model.

In this study, the pseudo-first-order, pseudo-second order], and intraparticle diffusion models were applied to observe the adsorption rates of BPA onto the adsorbent, as calculated by equation (7), (8), and (9), respectively [14].

$$\ln (q_e - q_t) = \ln q_e - k_1 t \tag{7}$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \tag{8}$$

$$q_t = k_{diff} t^{1/2} + C \tag{9}$$

### 2.6 Thermodynamic Study.

The thermodynamic studies were carried out at different temperature conditions, and calculated parameters included enthalpy ( $\Delta H$ ) and entropy ( $\Delta S$ ), as calculated by the following equation [15].

$$\ln K_c = \frac{\Delta S^\circ}{R} - \frac{\Delta H^\circ}{RT} \tag{10}$$

$$K_c = \frac{C_{e(adsorbent)}}{C_{e(solution)}} \tag{11}$$

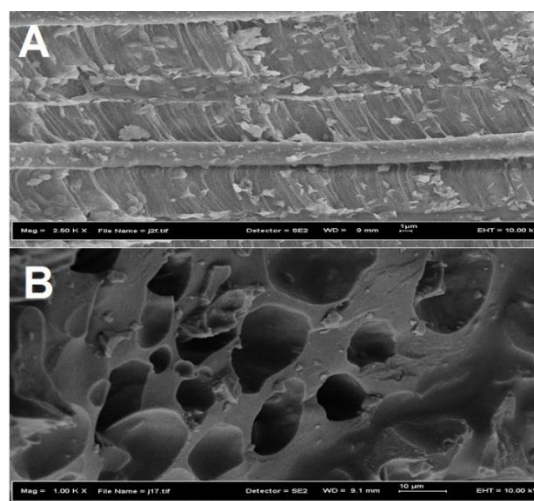
### 2.7. SEM and FTIR.

Morphological and surface functional group characteristics of the Banana bunch (*Musa acuminata*) and coconut bunch (*Cocos nucifera*) before and after adsorption were carried out using Scanning electron microscopy (JOEL-633F-SEM, Japan) and Fourier transform infrared (FTIR) spectroscopy (Spectrum one, Perkin Elmer, USA), respectively. The adsorbents were then characterized by FESEM, FTIR, and BET for before and after treatment of sulphuric acid.

## 3. Results and Discussion

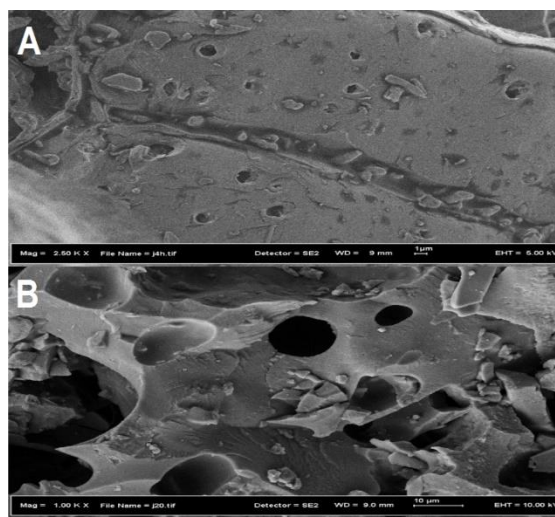
### 3.1. SEM, BET, and FTIR analysis.

Morphological characterization of the banana bunch and coconut bunch using SEM, before and after activation, is shown in Figure 2 and Figure 3. Prior to activation, the banana bunch showed more pore of different sizes and surface areas that actively react with BPA. The BET showed that the surface area of post-activation was increased from 4.80 m<sup>2</sup>/g to 5.83 m<sup>2</sup>/g, respectively.



**Figure 2.** Field Emission Scanning Electron Microscope (FESEM) image of the banana bunch: raw (A) and after activation (B).

This result indicated that the sulphuric acid treatment responsible for the growth in the porosity of the banana bunches, increase the number of available binding sites and its ion-adsorption capacity [16-18]. The post-activation morphology of the coconut bunch revealed a rough surface morphology indicating that the pore had been filled with BPA molecule, which implies a strong interaction between BPA and coconut bunch. However, the BET showed the depleting of surface area from 3.8 m<sup>2</sup>/g to 2.2 m<sup>2</sup>/g. The previous study showed that the pore volume, pore size distribution, functional groups of the contaminants, surface area, surface chemistry of adsorbent, and the solution characteristics played a significant role in the effectiveness of adsorbent to eliminate the sorbates [9,19,20].

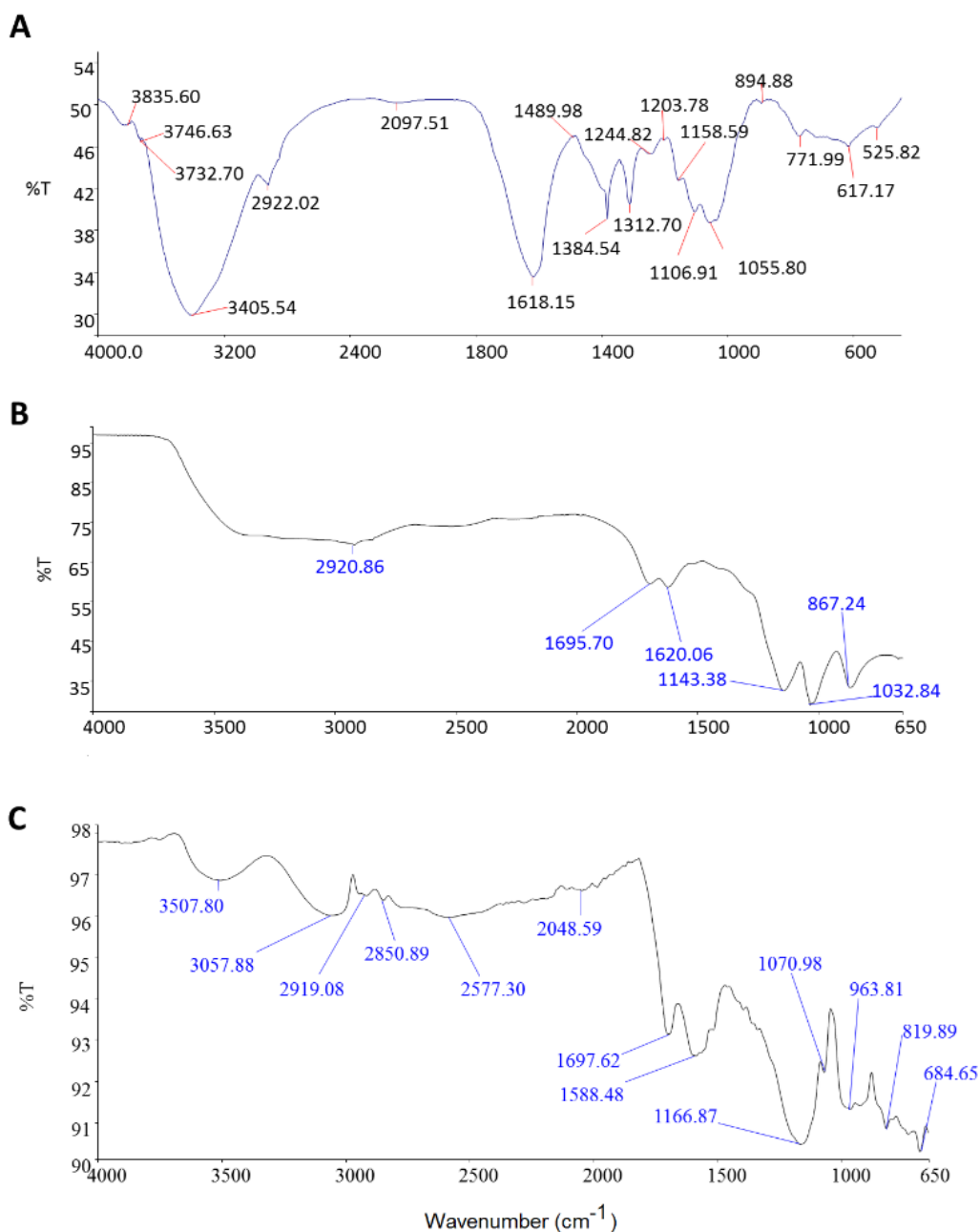


**Figure 3.** Field Emission Scanning Electron Microscope (FESEM) image of coconut bunch: raw (A) and after activation (B).

FTIR spectra of the banana bunch and coconut bunch before and after adsorption are shown in Figure 4 and Figure 5. From Fig. 4, it has seen the broad peak at 3405.54 cm<sup>-1</sup>, which indicated the involvement of hydroxyl groups, which possibly present in lignin, hemicellulose, and cellulose of the banana bunch. The oscillations of the N–H and O–H functional groups can be seen at the wavelength of 3405 cm<sup>-1</sup> and 2922 cm<sup>-1</sup>. Structure group of CH, CH<sub>2</sub>, and CH<sub>3</sub> contain in banana bunch also causes stretching of C–H which proved by a peak appeared at a wavelength of 2097 cm<sup>-1</sup>, 2922 cm<sup>-1</sup>, and the weak bands at 2920 cm<sup>-1</sup> while a peak that showed at a wavelength of 1618 cm<sup>-1</sup> indicated as C=C bonds in aromatic rings bending mode of the absorbed water. The peak showing at 1489 cm<sup>-1</sup> and 1244 cm<sup>-1</sup> known as the C–O carboxyl bands that indicated the activity of carboxyl oxygen atoms occurred. Besides, the figure showed CH<sub>2</sub> bends, OH bends, and C–O skeletal vibrations at the bands of 1489, 1384, and 1313 cm<sup>-1</sup>, respectively. While, C–O is stretching in hemicellulose at the band of 1244 cm<sup>-1</sup>.

This result also showed the presence of silica as bending peaks at 1143 cm<sup>-1</sup>, and 1032 cm<sup>-1</sup> indicated the Si–O stretching. The result of FTIR for the treated banana bunch showed an overlapped bands of sulphonic acid hydroxyl and free phenolic and alcoholic hydroxyl groups of the banana bunch (3400 cm<sup>-1</sup>). Besides, the presence of SO<sub>3</sub>H groups was definite by the peaks shifting at 1143 cm<sup>-1</sup> (SO<sub>2</sub>) and 1032 cm<sup>-1</sup> (SO<sub>2</sub>). The presence of carboxylic acid groups also was confirmed by peak shifting at bands of 1695 cm<sup>-1</sup> (C O) and 1620 cm<sup>-1</sup> (C O) [16]. FTIR spectrum of coconut bunch showed the possible involvement of the hydroxyl group as band shifting around the broad peak at 3406.36 cm<sup>-1</sup>. The functional group of CH, CH<sub>2</sub>, and CH<sub>3</sub> contain in coconut bunch also causes stretching of C–H, which proved by a peak

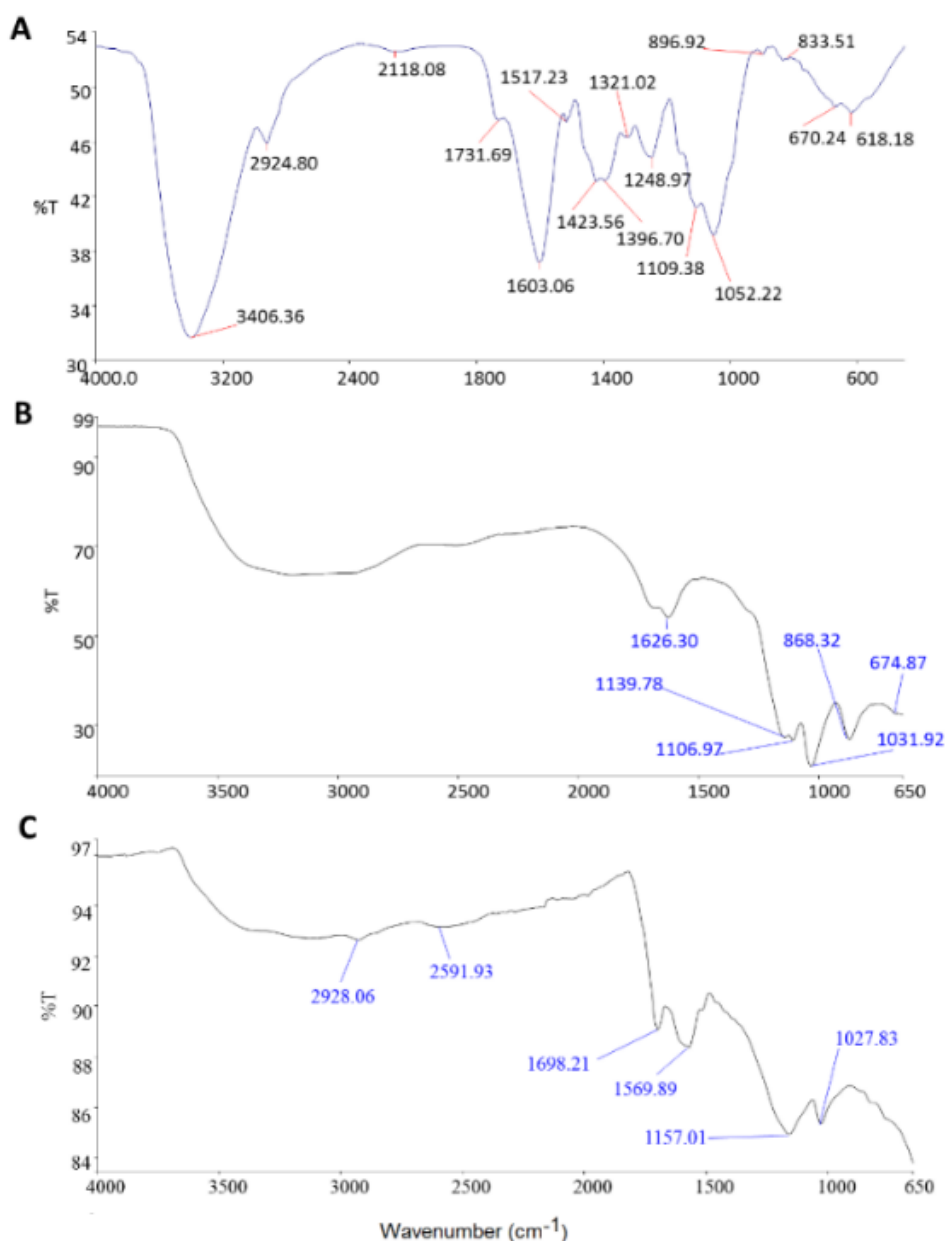
appeared at bands of  $2924\text{ cm}^{-1}$ . The vibrations of the N–H and O–H functional groups can be seen at the wavelength of  $3406\text{ cm}^{-1}$  and  $2118\text{ cm}^{-1}$ . The peaks at  $1731\text{ cm}^{-1}$  and  $1248\text{ cm}^{-1}$  indicated as C=C bonds in aromatic rings which also known as absorption band. The peak showing at  $1603\text{ cm}^{-1}$  and  $1052\text{ cm}^{-1}$ , known as the C–O carboxyl bands that indicated the activity of carboxyl oxygen atoms occurred.



**Figure 4.** FTIR spectrum of the raw banana bunch (A) and banana bunch after pre-treatment (B) and the banana bunch after sorption (C).

This result also showed the presence of silica as bending peaks at  $1139\text{ cm}^{-1}$ , and  $1031\text{ cm}^{-1}$  indicated the Si–O stretching band. The efficiency of the banana bunch and coconut bunch treatment is shown on the dominance of wavenumbers in inorganic groups. The presence of some functional groups and absorption peak in the adsorbent indicated the complex nature of the adsorbent from banana and coconut bunch. An increase and decrease in intensity of the peak occur to observe the role of the functional group of adsorbents in sorbates bonding. It was revealed that carbon and oxygen in the form of carboxylic acid was the most dominant element

in the adsorption mechanism. The presence of a sulfur element during chemical activation was also the main factor in increasing the elimination of contaminants [21,22].



**Figure 5.** FTIR spectrum of raw coconut bunch (A) and coconut bunch after pre-treatment (B) and the coconut bunch after sorption (C).

### 3.2. Batch studies.

Table 2 revealed the effect of parameters on the adsorption of BPA by the banana bunch and coconut bunch. The highest removal of BPA by the banana bunch and coconut bunch was obtained in the following condition: the temperature at 25° (75.53 and 77.70%), pH 3 (74.57 and 77.39%), dosage at 100 mg (75.53 and 77.71%), and agitation at 150rpm (76.77 and 77.12%). Previous studies reported that low pH value influenced the ionic state of the solution, thus increasing the adsorption process. Alkaline condition stimulates the electrostatic repulsion between BPA and the adsorbent, and changing the surface area of adsorbent becomes negative and inclined the binding affinity [23-26]. Thus, the optimum adsorbent dosage for this experiment was 100 mg indicating an advantage for economic purposes due to the usage of a

low quantity of adsorbent to reach high sorption of BPA [27, 28]. The adsorption was not impacted by agitation since there is no pattern of speed in increasing the BPA removal. In this case, moderate agitation (150 rpm) resulting in the highest removal sorption of BPA.

**Table 2.** Effect of Parameters on the adsorption of BPA.

Parameters	Adsorbents		
	Banana Bunch (%)	Coconut Bunch (%)	
Temperature (°C)	25	75.53	77.70
	30	72.42	74.91
	40	62.80	69.32
	50	55.03	64.66
	60	41.06	61.24
pH	3	74.57	77.39
	5	73.66	76.45
	7	72.11	75.22
	9	69.94	72.42
Adsorbent Dosage (mg)	50	69.32	55.65
	100	75.53	77.71
	300	60.93	66.52
	500	49.13	64.65
Agitation (rpm)	50	73.35	64.69
	100	75.84	52.90
	150	76.77	77.12
	200	75.53	60.23
	250	73.98	72.34

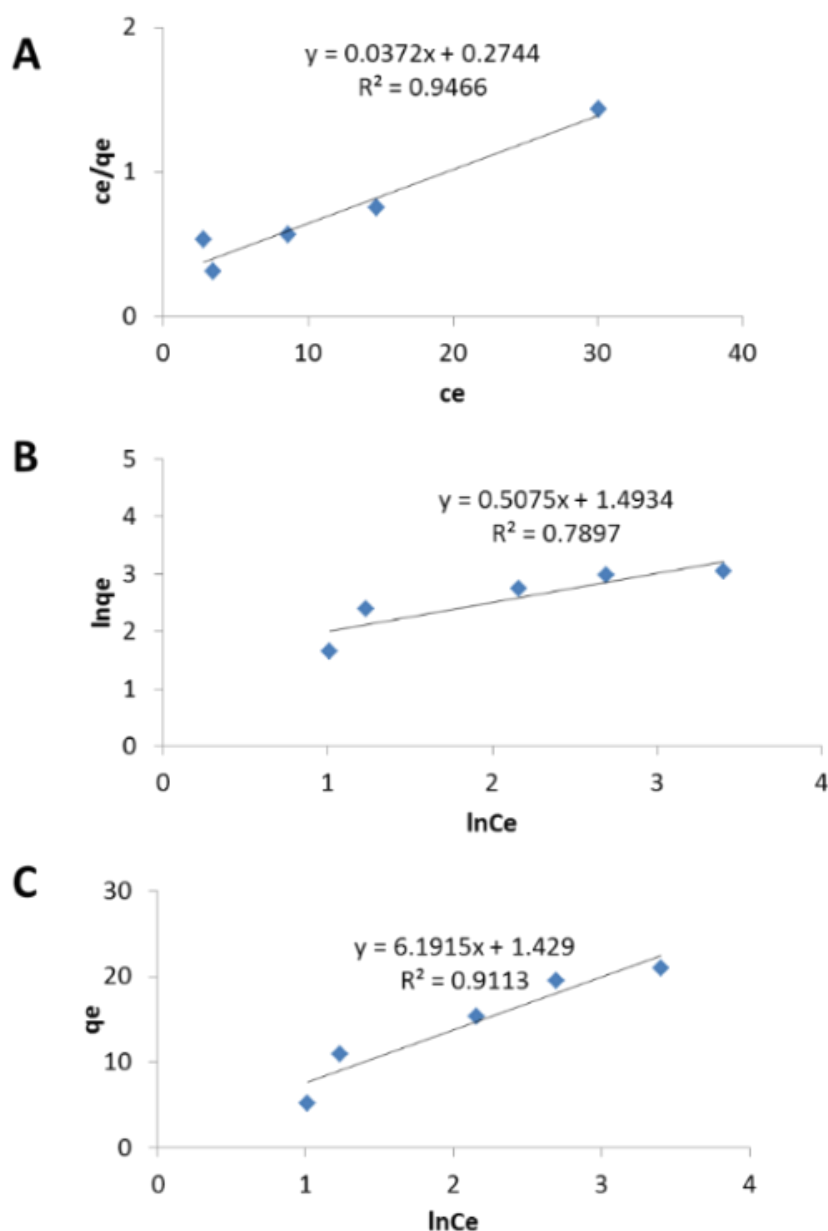
3.3. Sorption isotherm model.

**Table 3.** Effect of initial concentration on BPA adsorption onto banana bunch and coconut bunch.

C <sub>0</sub>	Removal (%)		q <sub>e</sub> (mg/g)	
	Banana bunch	Coconut bunch	Banana bunch	Coconut bunch
5	48.81	53.78	0.73	0.81
10	68.81	59.5	2.06	1.79
20	86.27	86.58	5.18	5.20
40	91.42	90.65	10.98	10.88
60	85.59	89.01	15.41	16.02
80	81.58	81.73	19.58	19.62
100	69.92	72.4	20.98	21.72

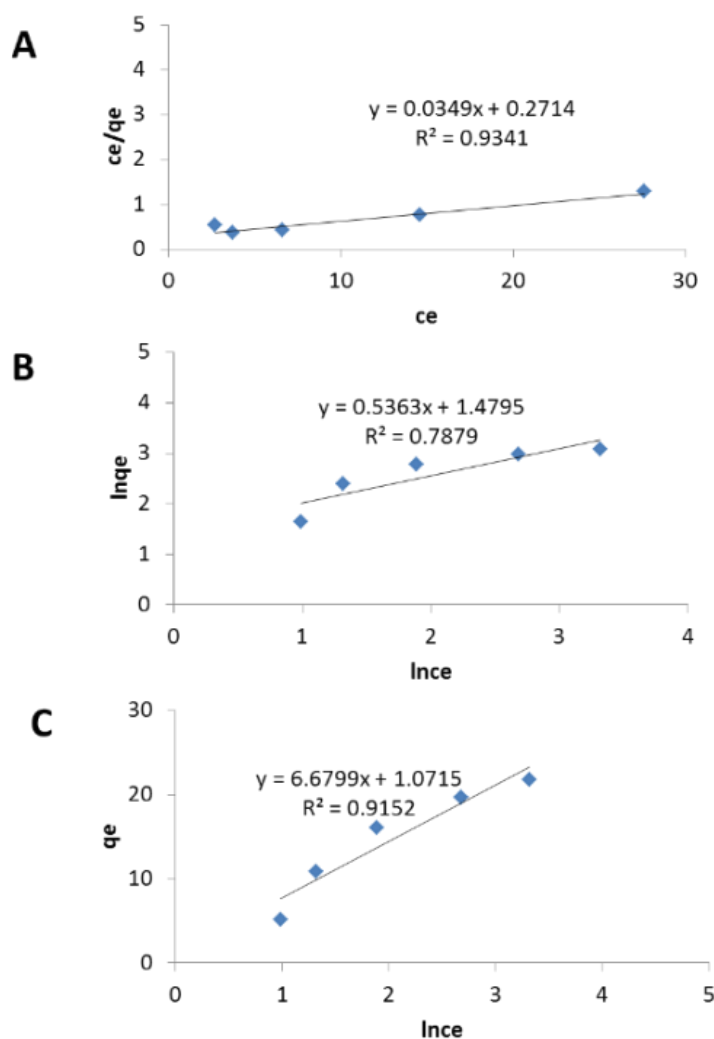
In order to evaluate the removal of BPA in aqueous solution and to understand its movement during the adsorption process, the sorption behavior and mechanism of BPA should be disclosed. The information about the equilibrium stage through the isotherm model is very important to know the relationship between adsorbent and sorbate. Table 3 reveals an increase in the equilibrium sorption capacity of the banana bunch (*Musa acuminata*) and coconut bunch (*Cocos nucifera*) with increasing initial BPA concentration. This current study showed that, at the initial stages, the adsorption capacity (q<sub>e</sub>) was rapidly increased and then gradually decreased after the optimum rate. This was suggested due to the ability of adsorbent per unit mass in adsorbing the accumulation of BPA. Therefore, the adsorption capacity of adsorbent was influenced by the initial concentration of BPA [29, 30]. As the concentration was increased from 5 to 100 mg/L, the adsorption capacity of BPA was also increased. This was due to the higher probability of collision between adsorbent particles and BPA ions. The higher initial concentration of BPA enhanced the adsorption capacity since the driving force was higher at the initial stage to overcome all mass transfer resistances of the BPA between the adsorbent and adsorbate [31-33]. While, for the removal rate of BPA, when the BPA concentration increased from 5 to 40 mg/L, the percentage removal was increased, while, when the initial concentration increased from 60 – 100 mg/L, the percentage removal decreased.





**Figure 6.** Adsorption isotherm of banana bunch; Langmuir (A), Freundlich (B) and Temkin (C).

Hence, the removal rate of BPA was dependent on the total available sorption sites of adsorbent, which were relatively limited and then resulting in a drop in the percentage removal of BPA when all the active sites were fully occupied [5, 34]. In this section, Langmuir, Freundlich, and Temkin were chosen as the most commonly used models for evaluating the experimentally acquired adsorption isotherms (Figures 6 and Figure 7). The values of the model parameters and the isotherm error deviation data for those model equations determined by linear regression analysis. The highest  $R^2$  for adsorption of BPA by was shown by the banana bunch, and the coconut bunch is Langmuir (0.947 and 0.934), indicating the monolayer coverage of pollutant molecule onto a solid surface. This model also reveals that once the active adsorbent site is completely filled with the adsorbate, the adsorption process would stop at that site [35-37]. However, the Langmuir adsorption model is only capable of explaining the adsorption capability and equilibrium process and not capable of determining the adsorption mechanism. [36,38].



**Figure 7.** Adsorption isotherm of coconut bunch; Langmuir (A), Freundlich (B), and Temkin (C).

### 3.4. Sorption kinetic model.

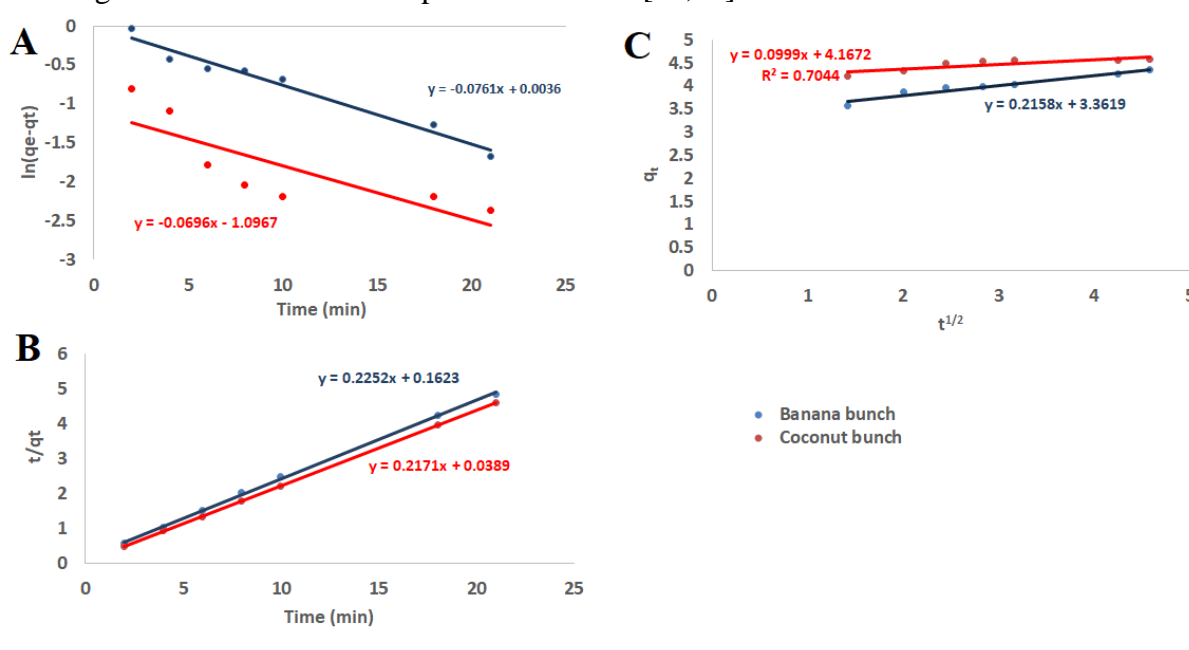
Investigation of the kinetic model is an important part of the adsorption process because it provides information about the reaction pathway and rate-controlling mechanism of the reaction. Table 4 shows the influence of the contact time on the BPA adsorption onto the banana bunch (*Musa acuminata*) and coconut bunch (*Cocos nucifera*).

**Table 4.** Effect of contact time on BPA adsorption onto banana bunch and coconut bunch.

Time (min)	Removal (%)		$q_e$ (mg/g)	
	Banana bunch	Coconut bunch	Banana bunch	Coconut bunch
2	59.38	70.29	3.563	4.215
4	64.66	72.11	3.880	4.327
6	65.9	74.91	3.954	4.494
8	66.21	75.53	3.973	4.532
10	67.14	75.84	4.029	4.550
18	70.9	75.84	4.252	4.550
21	72.42	76.15	4.345	4.569
24	75.53	77.7	4.532	4.662
30	75.22	77.4	4.513	4.643
36	75.22	76.7	4.513	4.606
38	74.91	76.68	4.494	4.588

Both banana and coconut bunch adsorbents showed that the optimum adsorption equilibrium was reached at the time of 24<sup>th</sup> min with the percentage removal and adsorption

capacity of BPA were 75.53%, 4.53 mg/g, and 77.7%, 4.66 mg/g respectively. The outcomes show that the removal rate and adsorption capacity increase gradually over time because, in the early stages of adsorption, there are a lot of empty sites and spread around the surface and easy to access, the impact of chemical activation on adsorbents so that absorption runs quickly in the beginning [26,39]. After equilibrium, the active adsorbent site was saturated with the sorbates causing the stagnancy of removal rate and adsorption capacity after they reach the highest value [5,40,41]. The results were also similar to some studies that, In the early phase of the adsorption process, the concentration of pollutants dissolved in the aqueous solution was very high, while the empty and active pore were dispersed on the surface of the adsorbent. Thus the removal rate and sorption capacity were rapidly increased. However, with increasing contact time, the number of active pores on the surface of the adsorbent was decreased, resulting in the decrease of adsorption of sorbates [29,42].



**Figure 8.** Adsorption kinetics plots for BPA using: (A) Pseudo-first order, (B) Pseudo-second order, and (C) intra-particle diffusion model.

**Table 5.** Adsorption capacity, isotherm and kinetic for BPA adsorption on different materials.

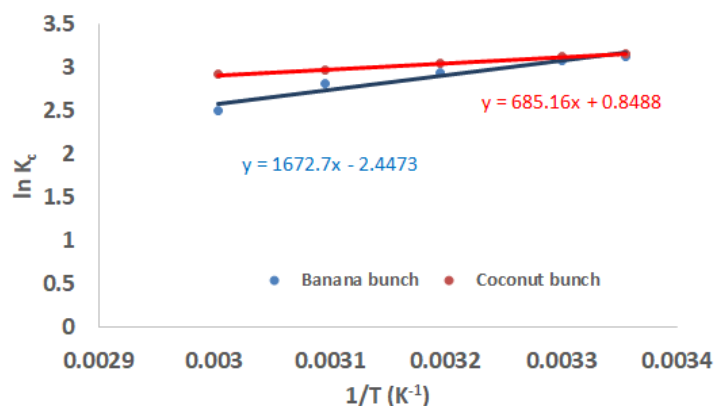
Adsorbent	Isotherm models	Kinetic Model	q <sub>e</sub> (mg/g)	References
Rice husk ash	Freundlich, Tempkin	Pseudo-second order	8.72	[44]
Commercial granular activated carbon	Freundlich, Tempkin	Pseudo-second order	3.54	[44]
Empty fruit bunch activated carbon	Langmuir	Pseudo-second order	41.98	[45]
Surfactant-modified natural zeolite	Tempkin	Pseudo-second order	6.9	[46]
Banana bunch	Langmuir	Pseudo-second order	4.53	This study
coconut	Langmuir	Pseudo-second order	4.66	This study

The Adsorption kinetics plots for BPA using Pseudo-first order, Pseudo -second order, and intra-particle diffusion model was shown in Figure 8. This study showed that the R<sup>2</sup> value for the pseudo-second-order model was found to be higher than that of the pseudo-first-order model and intraparticle diffusion model. The highest q<sub>e</sub> of the pseudo-second-order was closer to the adsorption of BPA at equilibrium (4.5 and 4.7 mg/g) than the value of another kinetic model. Even though the R<sup>2</sup> values of both models showed a good correlation, as it can be seen from the high R<sup>2</sup> values of 0.97 (the pseudo-first-order model) and 0.94 (intraparticle diffusion), respectively. In most reports focusing on BPA removal, the pseudo-second-order

fitted best to the kinetic adsorption data, as shown in Table 5. Thus, the adsorbents proposed that the overall rate of the BPA adsorption mechanism was chemisorption. This chemisorption mechanism involves the valence force of chemical bonds via electron exchange between the adsorbent and the pollutants [5,24,43].

### 3.5. Adsorption thermodynamics.

It is commonly known that the adsorption mechanism is strongly influenced by temperature, and these processes are controlled by thermodynamic parameters such as the standard enthalpy change ( $\Delta H^0$ , J mol<sup>-1</sup>), the standard entropy change ( $\Delta S^0$ , J mol<sup>-1</sup> K<sup>-1</sup>) and the standard free Gibbs energy change ( $\Delta G^0$ , J mol<sup>-1</sup>) of the adsorption processes. The values of  $\Delta S^0$  and  $\Delta H^0$  were calculated from the intercept and slope of the plot between  $\ln K_c$  Vs  $1/T$  for BPA concentration of 20 mg/L (Fig. 9). The negative values of  $\Delta H^0$  suggested that the adsorption of BPA onto banana bunch (-13.748 J mol<sup>-1</sup>) and coconut bunch (-5693.67 J mol<sup>-1</sup>) is exothermic in nature. Furthermore, the reduction of strong sorptive forces among the active sites in the adsorbent surface might cause a negative value of  $\Delta H^0$  [24,30,33]. The standard entropy change ( $\Delta S^0$ ) shows the negative value for a banana bunch (-19.84 J K<sup>-1</sup> mol<sup>-1</sup>) due to the decrease of randomness at the solid-solution interface at the time of adsorption of BPA from water body [30,33]. On the contrary, the coconut bunch shows the positive value (7.053 J K<sup>-1</sup> mol<sup>-1</sup>) due to the affinity of adsorbent in BPA removal and randomness at the solution interface during adsorption [30,33,47].



**Figure 9.** Thermodynamic adsorption of BPA by a banana bunch and coconut bunch.

## 4. Conclusions

The potential adsorbent derived from the banana bunch and coconut bunch were studied for BPA removal in aqueous solution. The highest removal of BPA by the banana bunch and coconut bunch was obtained in the following condition: the temperature at 25°, pH 3, dosage at 100 mg, and agitation at 150rpm. Alkaline condition stimulates the electrostatic repulsion between BPA and the adsorbent, and changing the surface area of adsorbent becomes negative and inclined the binding affinity. The highest R<sup>2</sup> for adsorption of BPA was shown by the banana bunch, and the coconut bunch is Langmuir (0.947 and 0.934), indicating the monolayer coverage of adsorbate molecule onto a solid surface. The adsorption of BPA at equilibrium (q<sub>e</sub>) for banana bunch was 4.53 mg/g, which is almost the same as the calculated value for the pseudo-second-order model, which is 4.50 mg/g. While the R<sup>2</sup> in the pseudo-second-order model gives a value of 0.99 that indicated a better fit, which is a little bit higher than the pseudo-

first-order and intraparticle diffusion with  $R^2$  value, it was 0.97 and 0.94. Thus, the adsorbents proposed that the overall rate of the BPA adsorption mechanism was chemisorption.

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## Conflicts of Interest

The authors declare no conflict of interest.

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